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CHE100 Stoichiometry Problems Instructor: S. Henzel  
 $2 \text{ C}_3\text{H}_6\text{O}_2 + 7 \text{ O}_2 \rightarrow 6 \text{ CO}_2 + 6 \text{ H}_2\text{O}$   
 1. How many moles of  $\text{CO}_2$  are produced when 8.652 moles of  $\text{C}_3\text{H}_6\text{O}_2$  react?  
 $8.652 \text{ mol C}_3\text{H}_6\text{O}_2 \times \frac{6 \text{ mol CO}_2}{2 \text{ mol C}_3\text{H}_6\text{O}_2} = 25.96 \text{ mol CO}_2$   
 2. How many moles of  $\text{H}_2\text{O}$  are produced when 50.75 g of  $\text{O}_2$  react?  
 $50.75 \text{ g O}_2 \times \frac{1 \text{ mol O}_2}{32 \text{ g O}_2} \times \frac{6 \text{ mol H}_2\text{O}}{7 \text{ mol O}_2} = 1.359 \text{ mol H}_2\text{O}$   
 3. How many moles of  $\text{O}_2$  are requir...

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### [CHAPTER 09 STOICHIOMETRY NOTES \(ANSWERS\)](#)

Unit 5: Chemical Equations and Reactions & Stoichiometry  
 Chemistry Chapter 9: Stoichiometry  
 9.1: Calculating Quantities in Reactions  
 Avogadro's Number: - a group of  $(6.022 \times 10^{23})$  molecules = 1 mole  
 Stoichiometry: - the calculation of quantities in a chemical reaction.  
 - the coefficients of various reactants and or products form mole ratios.  
 - these mole ratios hold for moles, molecules or atoms.  
 Mole Rat...

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